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## Stoichiometry Chapter Test A Answers Core Teaching

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Stoichiometry (page 2 of 2) 19. The mass of element X found in  
1.00 mole of each of four different compounds is 28.0 g, 42.0 g,  
56.0 g, and 70 g, respectively. The possible atomic weight of X is a.  
8.00 b. 14.0

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## Chapter 12 Stoichiometry Test A Answers

Answers: Moles and Stoichiometry Practice Problems 1) How many  
moles of sodium atoms correspond to  $1.56 \times 10^{21}$  atoms of sodium?  
 $1.56 \times 10^{21} \text{ atoms Na} \times 1 \text{ mol Na} = 2.59 \times 10^{-3} \text{ mol Na}$  236.022 x  
10 atoms Na 2) Determine the mass in grams of each of the  
following: a. 1.35 mol of Fe  $1.35 \text{ mol Fe} \times 55.845 \text{ g Fe} = 75.4 \text{ g Fe}$   
1 mol Fe b. 24.5 mol O

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## Chapter 9 Stoichiometry Test Answers

SOLUTION: You are asked to calculate the amount (in moles) of  
oxygen atoms in a given sample. You are given the mass of the  
oxygen sample. Use the molar mass of oxygen (1 mol O = 16.00 g)  
to create a conversion factor that converts mass (in grams) to  
amount (in mol) of oxygen.  $\frac{124 \text{ g O}}{16.00 \text{ g}} \times 1 \text{ mol O} = 7.75 \text{ mol}$

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## Chapter 3 Stoichiometry

It is readily available in pdf, ppt, word, rar, txt, kindle, and also zip.  
Answer Key Mole/Stoichiometry.Test.Review. Answer Key  
Mole/Stoichiometry.Test.Review 1.  $6.022 \times 10^{23}$  particles ( (atoms,  
(molecules) ( ( 2. 1mole (=6.022x10<sup>23</sup>particles ( ( 1mole=molar  
(mass (1mole=22.4L (3.

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